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is a branch of  
science which  
deals with the  
study of heat  
and temperature  
and their  
relation to



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aqueous  
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suggests, we  
introduce a  
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thermodynamics  
and use it for  
the study of

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solutions. We  
evaluate the  
quantities of  
higher order  
derivative than  
the normal  
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functions. We  
allow these  
higher  
derivative data  
speak for  
themselves

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without

resorting to any  
model system. We

thus elucidate

the molecular

processes in

solution,

(referred to in

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“mixing scheme

), to the depth

equal to, if not

deeper, than

that gained by



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scheme. The boundary between the adjacent regions is associated with an anomaly in the third derivatives of  $G$ . The loci of the anomalies in the temperature-composition field form the line sometimes

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referred as  
"Koga line". We  
then take  
advantage of the  
anomaly of a  
third derivative  
quantity of  
1-propanol in  
the ternary  
aqueous  
solution,  
1-propanol –  
sample species –  
H<sub>2</sub>O. We use its

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induced change  
as a probe of  
the effect of a  
sample species  
on H<sub>2</sub>O. In this  
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clarified what a  
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that concerns  
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is a network of  
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which  
composition  
changes are  
brought about  
either by mass  
transfer between  
phases or by  
chemical  
reaction within  
a phase, or by  
both.

Phase Diagrams  
and

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environmental  
engineering,  
geology, glass  
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other fields of  
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experts involved

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topics, but structurally the book introduces the thermal-fluid sciences.

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Chapter 5 -

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achievements. It  
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trigger a  
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has in effect  
gone astray. It  
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scientists in  
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are novel and  
call for a

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detailed  
reconsideration  
of the field.

The third  
chapter

acquaints the  
readers with  
their fourth  
colleague in  
Germany who was  
working on the  
proper link  
between the  
macroscopic

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thermodynamics, kinetics, and the atomistic representation of matter. The fourth chapter brings in their fifth colleague in the United States who could formally infer the famous formula  $S = k * \ln(W)$ ,

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companion to  
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for the  
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scientist. This  
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the range of

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practicing

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scientist who needs to apply thermodynamic relationships to aid in the understanding of the chemical process under consideration. The writing style in this volume remains informal, but more technical

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And its  
Principles and  
Applications. It  
starts with  
Chapter 11,  
which summarizes  
the  
thermodynamic  
relationships  
developed in  
this earlier  
volume. For  
those who want  
or need more

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detail, references are given to the sections in Principles and Applications where one could go to learn more about the development, limitations, and conditions where these equations apply. This is

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the only place  
where Advanced  
Applications  
ties back to the  
previous volume.  
Chapter 11 can  
serve as a  
review of the  
fundamental  
thermodynamic  
equations that  
are necessary  
for the more  
sophisticated

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Applications  
described in the  
remainder of  
this book. This  
may be all that  
is necessary for  
the practicing  
scientist who  
has been away  
from the field  
for some time  
and needs some  
review. The  
remainder of

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And this book  
applies  
thermodynamics  
to the  
description of a  
variety of  
problems. The  
topics covered  
are those that  
are probably of  
the most  
fundamental and  
broadest  
interest.

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Throughout the book, examples of "real" systems are used as much as possible. This is in contrast to many books where "generic" examples are used almost exclusively. A complete set of references to



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all sources of  
data and to  
supplementary  
reading sources  
is included.

Problems are  
given at the end  
of each chapter.  
This makes the  
book ideally  
suited for use  
as a textbook in  
an advanced  
topics course in

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chemical  
thermodynamics.  
An excellent  
review of  
thermodynamic  
principles and  
mathematical  
relationships  
along with  
references to  
the relevant  
sections in  
Principles and  
Applications

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where these  
equations are  
developed  
Applications of  
thermodynamics  
in a wide  
variety of  
chemical  
processes,  
including phase  
equilibria,  
chemical  
equilibrium,  
properties of

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mixtures, and  
surface  
chemistry Case-  
study approach  
to demonstrate  
the application  
of  
thermodynamics  
to biochemical,  
geochemical, and  
industrial  
processes  
Applications at  
the "cutting

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edge" of  
thermodynamics  
Examples and  
problems to  
assist in  
learning  
Includes a  
complete set of  
references to  
all literature  
sources

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## Applications Solutions Manual