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Calculating Moles

from Volume. Note

the Pattern;

Example 12.2.1;

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any stoichiometry

problem, the first

step is always to

calculate the

number of moles of

each reactant

present. In this

case, we are given

the mass of $K_2Cr_2O_7$

in 1 mL of

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solution, which we
can use to
calculate the
number of moles of
K₂Cr ...

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with oxygen to

produce aluminum
oxide as follows:



a. If you use

2.3 moles of Al,

how many moles of

Al_2O_3 can you

make? b. If you

want 3.9 moles of

Al_2O_3 , how many

moles of O_2 are

needed? 2. In the

presence of sulfuric

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acid, metallic iron
forms iron(III)
sulfate: $2\text{Fe} + 3\text{H}_2\text{SO}_4 \rightarrow \text{Fe}_2(\text{SO}_4)_3 + 3\text{H}_2$

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. What is stoichiometry based on? the law of conservation of mass. What does stoichiometry involve? balancing chemical equations and mole ratios.

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conversion factor
that relates the
number of moles of
any two ...

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