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Thermodynamics of Biochemical Reactions emphasizes the fundamental equations of thermodynamics and the application of these equations to systems of biochemical reactions. This emphasis leads to new thermodynamic potentials that provide criteria for spontaneous change and equilibrium under the conditions in a living cell.

Thermodynamics is the science that describes the behavior of matter at the macroscopic scale, and how this arises from individual molecules. As such, it is a subject of profound practical and fundamental importance to many science and engineering fields. Despite extremely varied applications ranging from nanomotors to cosmology, the core concepts of thermodynamics such as equilibrium and entropy are the same across all disciplines. A Conceptual Guide to Thermodynamics serves as a concise, conceptual and practical supplement to the major thermodynamics textbooks used in various fields. Presenting clear explanations of the core concepts, the book aims to improve fundamental understanding of the material, as well as homework and exam performance. Distinctive features include: Terminology and Notation Key: A universal translator that addresses the myriad of conventions, terminologies and notations found across the major thermodynamics texts. Content Maps: Specific references to each major thermodynamic text by section and page number for each new concept that is introduced. Helpful Hints and Don'ts: Try Its: Numerous useful tips for solving problems, as well as warnings of common student pitfalls. Unique Explanations: Conceptually clear, mathematically fair, simple, yet also sufficiently precise and rigorous. A more extensive set of reference materials, including older and newer editions of the major textbooks, as well as a number of less commonly used titles, is available online at <http://www.conceptualthermo.com/> 't Try Its: Numerous useful tips for solving problems, as well as warnings of common student pitfalls. Unique Explanations: Conceptually clear, mathematically fair, simple, yet also sufficiently precise and rigorous. A more extensive set of reference materials, including older and newer editions of the major textbooks, as well as a number of less commonly used titles, is available online at [http://www.conceptualthermo.com/a](http://www.conceptualthermo.com/). Undergraduate and graduate students of chemistry, physics, engineering, geosciences and biological sciences will benefit from this book, as will students preparing for graduate school entrance exams and MCATs.

The Instructor's solutions manual to accompany Atkins' Physical Chemistry provides detailed solutions to the 'b' exercises and the even-numbered discussion questions and problems that feature in the ninth edition of Atkins' Physical Chemistry. The manual is intended for instructors and consists of material that is not available to undergraduates. The manual is free to all adopters of the main text.

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