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 Answer Key Chapter 9 - Chemistry 2e | OpenStax. 1. The cutting edge of a knife that has been sharpened has a smaller surface area than a dull knife. Since pressure is force per unit area, a sharp knife will exert a higher pressure with the same amount of force and cut through material more effectively. 3.**

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Chemistry Chapter 9 Answers. Flashcard maker : Lily Taylor. How can you determine the charges of monatomic ions? When the metals of group 1A, 2A, and 3A lose electrons, the form cations with positive charges equal to their group number.

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zombiekiller_sarah PLUS. CHAPTER 9-Mastering Chemistry. d. strong electrolyte. c. weak electrolyte. e. the amount of ion that carries 1 mol-. e. saturated. In water, a substance that ionizes completely in solution is c-. In water, a substance that partially ionizes in solution is ca-. An equivalent is ...

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9.1. Write the formulas for the following coordination compounds: (i)Tetraamminediaquacobalt (III) chloride. (ii)Potassium tetracyanonickelate (II) (iii)Tris (ethanp-1,2-diamine) chromium (III) chloride. (iv)Amminebromidochloridonitrito-N- platinatc (II) (v)Dichloridobis (ethane-1 ,2-diamine) platinum (IV) nitrate.

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This chapter is all about the solution. Solution Chemistry class 9 PDF notes about the water solution. There are different types of solutions but Saturated and supersaturated solutions are the most important questions. The Tyndall effect, suspension, collide are important for MCQs, and short questions.

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Chemistry (4th Edition) Burdge, Julia Publisher McGraw-Hill Publishing Company ISBN 978-0-07802-152-7

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Answer: (a) Primary valency is ion isable. Question 6. Ammonia acts as a very good ligand but ammonium ion does not form complexes because. (a) NH 3 isa gas while NH is in liquid form. (b) NH 3 undergoes sp 3 hybridisatiort while NH undergoes sp 3 d hybridisation. (c) NH ion does not have any lone pair of electrons.

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As we noted in Chapter 9.2 the heat released by a reaction carried out at constant volume is identical to the change in internal energy (ΔE) rather than the enthalpy change (ΔH); ΔE is related to ΔH by an expression that depends on the change in the number of moles of gas during the reaction. The difference between the heat flow measured at constant volume and the enthalpy change is usually quite small, however (on the order of a few percent).

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