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Equations OBJECTIVES: Explain how  
balanced equations apply to both  
chemistry and everyday life. Section  
12.1 The Arithmetic of Equations  
OBJECTIVES: Interpret balanced  
chemical equations in terms of: a)

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moles, b) representative particles, c) mass, and d) gas volume (Liters) at STP.

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Arithmetic of Equations 12.2

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Balance the following equations and  
then solve the related problems: 1)

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An expanded version of the flowchart for stoichiometric calculations illustrated in Figure 11.4.1 is shown in Figure 12.2.1. We can use the balanced chemical equation for the reaction and either the masses of solid

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reactants and products or the volumes of solutions of reactants and products to determine the amounts of other species, as illustrated in Example 7, Example 8, and Example 9.

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PLAY. . . . Use the following balanced equation to answer the question:  $\text{Mg} + 2\text{H}_2\text{O} \rightarrow \text{Mg}(\text{OH})_2 + \text{H}_2$  If 16

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chapter 12. 2 Mg are heated with 12.



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0g H<sub>2</sub>O, which is the limiting  
reactant and which is the . . .

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Example 12.2.1 and Example 12.2.2,  
the identity of the limiting reactant  
has been apparent:  $[\text{Au}(\text{CN})_2]^-$ ,  $\text{LaCl}_3$ , ethanol, and para-nitrophenol.

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