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## Chapter 11

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Chemistry: chemical  
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TEACHER GUIDE AND ANSWERS

Study Guide - Chapter 11 -  
Stoichiometry Section 11.1

What is stoichiometry? 1.

true 2. true 3. false 4.

true 5. true 6. 2, 2, 64.10

7. 3, 3, 96.00 8. 2, 2,

88.02 9. 4, 4, 72.08 10.

methanol and oxygen gas 11.

carbon dioxide and water 12.

160.10 g 13. 160.10 g 14.

They are equal. 15. A mole

ratio is a ratio between the  
numbers of moles

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### Technology

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Stoichiometry Section

11.1.1 Objectives Describe the types of relationships indicated by a balanced chemical equation. State the mole ratios from a balanced chemical equation. Review Vocabulary reactant: the starting substance in a chemical reaction New Vocabulary stoichiometry mole ratio Defining Stoichiometry

### **Chapter 11: Stoichiometry**

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Li. 11.1 Defining  
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**Chemistry Matter and Change:  
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CHAPTER STUDY GUIDE Date  
Class Stoichiometry Section  
11.1 What is stoichiometry?

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**Answer Key**  
In your textbook, read about stoichiometry and the balanced equation. For each statement below, write true or false. 2. 3. 4. 5. The study of the quantitative relationships between the amounts of reactants used and the amounts of products formed by a chemical reaction

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Defining Stoichiometry Vocab

- stoichiometry • mole ratio

Concepts Using Balanced

Equations • Number of Atoms

- Number of Molecules •

Number of Moles

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In Section 11.3 , for  
example, you learned how to  
express the stoichiometry of  
the reaction for the

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## Answer Key

ammonium dichromate volcano in terms of the atoms, ions, or molecules involved and the numbers of moles, grams, and formula units of each (recognizing, for instance, that 1 mol of ammonium dichromate produces 4 mol of water). This section describes how to use the stoichiometry of a reaction to answer questions like the following: How much oxygen is needed to ensure complete combustion of a ...

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The **Overview** section summarizes the content of the chapter and includes a comprehensive listing of terms, a summary of general concepts, and a list of numerical exercises, while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook. The Fill-in, Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples. The

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Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry, while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter. After briefly dealing with an overview of chemistry, this book goes on exploring the concept of matter, energy, measurement, problem solving, atom, periodic table, and chemical bonding. These topics are followed by discussions on writing names and formulas of compounds; chemical

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**Answer Key** formulas and the mole; chemical reactions; calculations based on equations; gases; and the properties of a liquid. The remaining chapters examine the solutions; acids; bases; salts; oxidation-reduction reactions; electrochemistry; chemical kinetics and equilibrium; and nuclear, organic, and biological chemistry. This study guide will be of great value to chemistry teachers and students.

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