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Impossible to
Climb 15 Meters
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diprotic acid
requires 42.57

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42 mL of 0.111 M

NaOH to
completely
neutralize a

0.685-gram
sample.

Calculate the
approximate
molar mass of
the acid. a. 423
g/mol b. 212
g/mol c. 290
g/mol...

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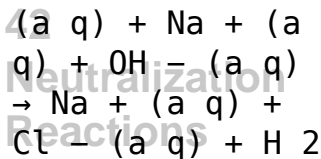
The full ionic equation for the neutralization of hydrochloric acid by sodium hydroxide is written as

follows: (4.4.2)

$$\text{H}^+ + (\text{a q}) + \text{Cl}^-$$

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(l) Since the acid and base are both strong, they are fully ionized and so are written as ions, as is the NaCl formed as a product.

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To determine the amount of acid and conjugate base in solution after the neutralization reaction, we calculate the

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42 amount of $\text{CH}_3\text{CO}_2\text{H}$

in the original

solution and the

amount of OH^-

in the NaOH

solution that

was added. The

acetic acid

solution

contained 50.00

mL (0.100 mmol ($\text{CH}_3\text{CO}_2\text{H}$) / mL) =

5.00 mmol ($\text{CH}_3\text{CO}_2\text{H}$)

5.00 mmol ($\text{CH}_3\text{CO}_2\text{H}$)

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42) The NaOH

solution

contained

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Reactions and

Titration Curves

...

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42 my guide. The balancing the whole equation video is a good example of how I support student thinking during this part of the lesson. The big challenge is around writing the chemical formulas for salts,

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42 especially for
salts containing
polyatomic ions.

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2 group 3 names.

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determination of

Enthalpy of

Neutralization

between HCl and

NaOH (Ccal:

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17.55 J/°C).

Calculation

example Trial 1

Trial 2 Trial 3

99.9 ml 99.9 ml

Volume of HCl

99.9 mL 24.0 °C

23.0 °C 25.0 °C

Initial

temperature of

HCl 99.8 ml 99.8

ml Volume of

NaOH 99.8 ml

23.0 °C 25.0 °C

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24.0 °C Initial
temperature of
NaOH 38.0 °C
40.0 °C 39.0 °C
mol mol mol
Final solution
temperature ...

Lab Notebook
Etermination Of
Enthalpy Of
Neutraliz ...

Let's use a
calorimeter to

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42 measure the energy changes accompanying neutralization reactions. You need this equipment: 2 Styrofoam (polystyrene) cups, cardboard square with hole in center, split one-hole rubber stopper,

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42 thermometer,
400-mL beaker,
Neutralization
Reactions
250-mL beaker,
50-mL graduated
cylinder. You
need these
materials: 1 M
HCl, 1 M HC 2 H
3 O 2 (acetic
acid), 1 M NaOH

*Heat of
Neutralization -
high school*

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Neutralization
Reactions
2. a) Based on
your

observations and
the answers to
post-lab
question #1
above, which
reactant was
initially
present in
excess amount,
the copper (II)

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42 chloride or the
aluminium? b)
Neutralization
Reactions
How does this
observation
compare with
your Pre-Lab
Prediction? If
possible,
propose an
explanation.
Carry out
calculations to
support this.

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42 *Acids and Bases:
Titration #1
Neutralization
Reactions*
*Determination of
[NaOH] by ...*

Hess' Law Lab

Practice

Worksheet - Heat
of

neutralization
of NaOH(aq) and
HCl(aq) 1. Write
molecular
equation and
NIE. 2.

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4.2 Calculate AT a.
Find T_{final} b.
Find average
 T_{initial} of
NaOH, find
average T_{initial}
of HCl, and
average T_{initial}
Overall. C. Find
AT 1. Calculate
heat of neutrali-
zation/mole of
water formed.

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Worksheet - Heat
Of Neutral ...

No. Name Matrix

No. Signature 1

ARIF HANAFI BIN

MD ZAKI

2016800892 2

MUHAMMAD

AMERUDDIN DANIEL

BIN SULONG

2016800832

Instructor 1..

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RAFIE BIN

DERAMAN 2.

Objective : To
erect the
calorimeter and
determine heat
capacity of a
calorimeter then
measure.

CHE 142

Inorganic

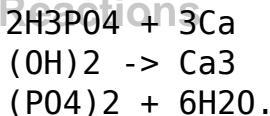
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Neutralization



Neutralization:

In many cases, the reactants are deadly when separate, but together, they are harmless (as water and salt).

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42 Acids release protons when dissolved in water, which allows it to react with metals. HCl , H_2SO_4 , H_2CO_3 .

*Neutralization
Reactions on
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Neutralization

Reactions require

more time to

spend to go to

the book

instigation as

with ease as

search for them.

In some cases,

you likewise get

not discover the

broadcast

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Neutralization
Reactions

neutralization reactions that you are looking for. It will extremely squander the time. However below, taking into ...

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Neutralization
Lab Answer Key

Exp #5 Heat

Neutralization

Answer: 50.0 g).

Add the mass of
HCl and the mass
of NaOH to give
the total mass
used, this will
be the mass you
will use to

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42 calculate heat
of reaction, q .

- Specific heat:

The specific

heat for

reaction 1 can

be assumed to be

close to that of

pure water

($4.184 \text{ J/g} \cdot ^\circ\text{C}$).

- ΔT :

Heat Of

Neutralization

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