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Neutralization reactions are one type of chemical reaction that proceeds even if one reactant is not in the aqueous phase. For example, the chemical reaction between $\text{HCl}(\text{aq})$ and $\text{Fe}(\text{OH})_3(\text{s})$ still proceeds according to the equation. $3 \text{HCl}(\text{aq}) + \text{Fe}(\text{OH})_3(\text{s}) \rightarrow 3 \text{H}_2\text{O}(\text{l}) + \text{FeCl}_3(\text{aq})$ even though $\text{Fe}(\text{OH})_3$ is not soluble.

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A neutralization reaction is a reaction in which an acid and a base react in an aqueous solution to produce a salt and water. The aqueous sodium chloride that is produced in the reaction is called a salt. A salt is an ionic compound composed of a cation from a base and an anion from an acid.

~~4.4: Neutralization Reactions — Chemistry LibreTexts~~

Read Chapter 19, pages 617 - 621 in the Glencoe - Chemistry Matter & Change textbook. Read Chapter 4, pages 149 - 154 (Section 4.8) in the Zumdahl - Chemistry textbook. Complete the "Neutralization Reactions: Titrations" booklet. (File and answer key below)

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the worksheet is actually very simple. All you need to do is to use a spreadsheet program and make the necessary adjustments to create the neutralization reaction chart.

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About This Quiz & Worksheet. This quiz and worksheet can help you gauge your knowledge of neutralization reactions and their characteristics. You will be assessed on acidic and basic substances.

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